

Assay of Isonicotinic Acid Hydrazide

Practical 07 | Medicinal Chemistry III

Aim

To assay Isonicotinic Acid Hydrazide, and interpret the recorded observations.

Principle

- Isoniazid is an isonicotinic-acid hydrazide. The assigned assay may use its defined redox or other analytical response; calculate with the exact reaction and validated method factor.
- The property is obtained from a defined observable under controlled conditions.
- Replicate measurements, instrument checks and consistent units are needed before comparison.

Materials and Apparatus

- Assigned isoniazid, reference or standardized analytical reagents and completed/observed assay readings.
- Assigned test material and calibration or comparison material.
- The apparatus specified for this property, a balance or volumetric equipment where relevant, and a record sheet.

Method

1. Identify the exact analyte form, assay method and reporting basis assigned to the exercise.
2. Record the standardized reagent factor or the reference-standard potency used.
3. Prepare the measured test portion and comparison/blank according to that method.
4. Obtain replicate endpoint or instrumental responses and retain the raw readings.
5. Apply reaction stoichiometry, reference potency and dilution corrections as applicable.
6. Report content on the specified mass or potency basis and compare only with the assigned criterion.

Observation Record

Complete this record using actual laboratory findings or the provided teaching data.

Observation / parameter	Record / finding
Analyte form	_____
Test portion	_____
Method and reference/factor	_____
Raw assay response	_____
Dilution	_____
Assay result	_____

Calculations and Interpretation

- Titrimetric method: $\text{analyte mass} = \text{corrected volume} \times \text{verified method factor} \times \text{dilution correction}$.
- Instrumental method: obtain concentration from the valid reference calibration and correct for preparation/dilution.
- Assay (%) = $\text{measured analyte amount} / \text{nominal analyte amount} \times 100$ on equivalent units and form.

Result

Record the actual finding, calculated value with units where relevant, and a conclusion supported by the observations.

Finding: _____

Conclusion: _____

Precautions

- Keep temperature and test conditions consistent across comparative readings.
- Record raw readings before rounding calculated values.
- Do not compare results obtained with different methods as if they were interchangeable.

Viva Questions

1. Why is the analyte form part of the calculation?

Salt, hydrate and declared potency conventions can change the factor and reporting basis.

2. Why control the test conditions?

A change in conditions may alter the measured property independently of the intended variable.

3. Why retain replicate readings?

They show repeatability and help identify an anomalous measurement.

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